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Magnesium chloride – ethanol adducts

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Some 60 addition compounds of the type $\text{MgCl}_2 \cdot n\text{EtOH}$ ($6 \leq n \leq 1.4$), formed between magnesium chloride and ethanol, have been prepared and were characterized by means of X-ray powder diffraction. The spectra were interpreted on the basis of a subset of nine characteristic diffraction patterns, and a range of (partially) new products, $\text{MgCl}_2 \cdot n\text{EtOH}$ with $n=6, 4.5, 4, 3.33, 2.5, 1.67, 1.50$ and 1.25 , is described. MgCl_2 adducts with alcohols can be converted to active olefin polymerization catalyst supports through the elimination of the alcohol molecules from the adducts by thermal desolvation.

1. Introduction

The drying process of bischofite, $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$, to obtain pure anhydrous MgCl_2 , has been given ample attention [1–7] and it has been shown [8] that by chlorination of partially dried bischofite with HCl gas at 400°C , a practically pure and dry MgCl_2 is obtained with only traces of MgO and H_2O . Efficient removal of water is also achieved by reaction with silanes [9] or by utilizing a magnesium chloride solvent, such as an alcohol, in a distillation zone whereby an azeotropic mixture of alcohol and water is removed overhead [2]. For the purpose of preparation of highly disperse (active) MgCl_2 -based supports for polyolefin Ziegler–Natta catalysts, there is considerable interest in the properties of MgCl_2 -electron-donor (Lewis-base) complexes [10–13], such as $\text{MgCl}_2 \cdot n\text{THF}$ [14, 15], $\text{MgCl}_2 \cdot n\text{EB}$ [16–19], $\text{MgCl}_2 \cdot n\text{EtOAc}$ [17, 19] and $\text{MgCl}_2 \cdot 2\text{B}$ [19] (THF, tetrahydrofuran; EB, ethylbenzoate; EtOAc, ethylacetate; B, monoesters, alcohols and ketones). Thermal or extractive desolvation of MgCl_2 -Lewis-base adducts provides very active supports. In the case of $\text{MgCl}_2 \cdot 4\text{THF}$, this process appears to proceed as follows: $\text{MgCl}_2 \cdot 4\text{THF} \rightarrow \text{MgCl}_2 \cdot 2\text{THF} \rightarrow \text{MgCl}_2 \cdot 1.5\text{THF} \rightarrow \text{MgCl}_2 \cdot \text{THF} \rightarrow \beta\text{MgCl}_2$ [15]. It is considered to be important that the interaction between MgCl_2 and the solvate is weak so that easy removal of these moieties (chemical activation) may form a very disordered MgCl_2 material, suitable as a Ziegler–Natta support [20].

Magnesium chloride-supported catalysts are increasingly made from magnesium chloride alcoholates (Mitsui–Himont technology) or by reacting magnesium alkyls and alkoxides with other halogenated materials. Procatalysts for the (co)polymerization of ethylene or propylene (and α -olefin), obtained by alcoholation of MgCl_2 powder, have been described [9–12, 21, 22]. On the other hand, magnesium chloride supports may also be formed by spray-drying ethanol solutions of the salt [11, 23]. Recently, various

aspects of the thermal desolvation of MgCl_2 adducts with alcohols and ethers were reported [20, 24]. The results of complete desolvation of $\text{MgCl}_2 \cdot 6\text{ROH}$ are controversial. While it has been stated [25] that free MgCl_2 is obtained by desolvation of $\text{MgCl}_2 \cdot 6\text{ROH}$, Multani [26] obtained ClMgOEt by heating $\text{MgCl}_2 \cdot \text{EtOH}$ to 180°C at $p = 0.1$ mm. Other workers have reported alkoxymagnesium chlorides as being the reaction products [24, 27]. We have also shown interest in the dealcoholation of ethanol associates of magnesium chloride. In order to gain insight into the parameters controlling this process, it was felt to be worthwhile to examine the MgCl_2 -ethylalcohol adduct compositions. Some 60 $\text{MgCl}_2 \cdot n\text{EtOH}$ solvates ($6 \leq n < 1.4$) were synthesized using various preparative routes.

2. Experimental procedure

$\text{MgCl}_2 \cdot n\text{EtOH}$ ($6 \leq n < 1.4$) adducts were prepared as white powders according to one of the synthetic routes detailed below.

2.1. Liquid-phase quenching of fused emulsified adducts (controlled turbulence emulsion method, CTEM)

Commercial MgCl_2 and a calculated amount of dry EtOH were loaded in an autoclave together with excess of paraffin oil (ROL–OB/30) and heated at 120 – 150°C for 1–3 h under continuous stirring in a nitrogen atmosphere. The emulsion was then ejected from the reaction vessel through a tube (length, 1.2 m; diameter 2 mm) with a 1 mm exit by means of a nitrogen stream with a flow rate of about 4 – 5 ms^{-1} at a constant pressure of 10 atm and the liquid phase was quenched in stirred n -heptane kept at -30°C . The suspension was held at this temperature for 1 h under stirring and after decanting for 15 min a microspheroidal product with mean particle size in the

30–50 μm range was collected in a nitrogen atmosphere. The solid was repeatedly washed with *n*-heptane to remove paraffin oil, filtered, dried at room temperature in vacuum and subjected to ultrasonics for further particle segregation.

2.2. Fluid-bed dealcoholation

$\text{MgCl}_2 \cdot n\text{EtOH}$, obtained according to preparative methods given in Sections 2.1, 2.5 or 2.7, was packed into a fluidized bed reactor (length, l 270 mm; diameter 25 mm) and evacuated to 0.1 torr (1 torr = 133.322 Pa) isothermally at temperatures ranging from 40–90 °C. A nitrogen-stream (flow rate 21 h^{-1} , corresponding to 4–5 mm Hg) was then passed through the packed reactor maintained at this temperature and calculated amounts of EtOH were collected in cold traps. After 15 min, about 2 cm^3 were collected from $5.6\text{ MgCl}_2 \cdot 6\text{EtOH}$, evacuated at 60 °C, indicating the formation of $\text{MgCl}_2 \cdot 2.85\text{EtOH}$. The method was equally well used as a post-treatment of $\text{MgCl}_2 \cdot n\text{EtOH}$ ($n < 6$) adducts prepared according to the aforementioned different experimental procedures.

2.3. Azeotropic distillation

$\text{MgCl}_2 \cdot 6\text{EtOH}$ was prepared at room temperature by slow evaporation under exclusion of humidity of a highly saturated MgCl_2 solution in ethanol. The product (23 g), obtained as (poor quality) colourless crystals, which are efflorescent and hygroscopic, was subjected to azeotropic distillation with 0.51 dry *n*-hexane in a 30 cm column packed with Fenske rings. The distillation was continued at total reflux until a constant temperature was achieved at the top of the column. The product was filtered in a nitrogen stream and vacuum dried at room temperature.

2.4. Vacuum-drying

$\text{MgCl}_2 \cdot 6\text{EtOH}$, obtained according to Di Noto *et al.* [20], was vacuum dried in a flask at 0.1–0.2 torr on a thermostatic bath at 60 °C for the time necessary to reach a desired composition. After 1 h outgassing, $\text{MgCl}_2 \cdot 3.85\text{EtOH}$ was obtained.

2.5. Recrystallization

$\text{MgCl}_2 \cdot n\text{EtOH}$ was prepared by adding n mol ethanol to 1 mol commercial anhydrous MgCl_2 under cooling to 0 °C in nitrogen atmosphere. Subsequently, the reaction tube was closed and heated until fusion of the contents under magnetic stirring. After cooling to room temperature, the desired product was obtained.

2.6. Reaction with HCl

A 80–125 μm fraction of metallic magnesium (C. Erba) was dissolved in *n*-heptane containing a stoichiometric amount of $\text{C}_2\text{H}_5\text{OH}$ and anhydrous HCl was passed through the mixture for 4 h at $T < 15^\circ\text{C}$. After

additional heating at 70 °C for 3 h in a stream of HCl, a white product was isolated.

2.7. Spray drying

MgCl_2 and dry EtOH were loaded in an autoclave at 12 atm and melted by heating at 120 °C for 1–3 h under continuous stirring. The product was then ejected from the reaction vessel through a tube (length, l 2 m; diameter 2 mm) with a 1 mm exit by means of a nitrogen stream at a pressure of 10 atm and the liquid phase was spray-cooled (by Niro Atomizer equipment) at 15–50 °C in a turbulent nitrogen atmosphere. The adducts were obtained in microspheroidal form with average particle diameter less than about 40 μm .

The following standard analytical techniques were used to verify the chemical composition of the white reaction products: magnesium, complexation with EDTA using Eriocromo T as an indicator; chlorine, potentiometric methods; EtOH [28]; H_2O , Karl Fischer.

X-ray powder spectra were recorded on a Philips PW 1050/1070 diffractometer using nickel-filtered CuK_α radiation ($\lambda = 0.15418\text{ nm}$) at 40 kV, 20 mA. Like most $\text{MgCl}_2 \cdot n\text{ROH}$ adducts, the $\text{MgCl}_2 \cdot n\text{EtOH}$ complexes are unstable, hygroscopic, deliquescent in an atmosphere saturated with alcohol vapour, and easily loose ethanol in an atmosphere in which the partial pressure is lower than the saturated vapour pressure. Therefore, the specimens were protected from the atmosphere by Nujol and were packed as a slurry into the X-ray diffractometer holder (paraffin oil shows up in the X-ray spectra as a broad band centred at 18° in 2θ). The characteristic d -values and peak intensities of the various $\text{MgCl}_2 \cdot n\text{EtOH}$ phases identified, on the basis of which the 60 XRD spectra could be interpreted, were obtained from diffractometer charts recorded with a scanning speed of $1/4^\circ$ and $1^\circ 20\text{ min}^{-1}$, after careful calibration of the goniometer with $\alpha\text{-Al}_2\text{O}_3$. Given the instability of the products, great care was taken in ensuring the reproducibility of both diffraction positions and intensities. Decomposition and transformation temperatures were recorded on a Mettler TA2 thermobalance equipped with a low-temperature furnace using $\alpha\text{-Al}_2\text{O}_3$ (DTA) as a reference sample. Temperatures were measured using a Pt/Pt–Rh thermocouple and heating rates of $1\text{--}3^\circ\text{C min}^{-1}$ were used. Heating curves were run in controlled atmospheric conditions.

3. Results and discussion

3.1. Solvation

The solvation behaviour of magnesium chloride in mixed aqueous organic solvents (amongst which is EtOH) has widely been studied and sets the background scene. Whereas MgCl_2 in the solid state consists of the hard Lewis-acid Mg^{2+} coordinated to six neighbouring Cl^- ligands in a typical layer structure (Fig. 1), in aqueous solution preferentially the more stable complexes with the harder H_2O base are expected [29, 30]. If we consider the order of decreasing

electronegativity, understood that ions may easily other factors, su a role in determ

In $\text{MgCl}_2/\text{H}_2\text{O}$ complexes of the strated by diffra molecular dynamics [39]. Raman [40] ance (NMR) spe number of eight (local) quasi-crys close-packing in [34, 37]. In very ions in water, ha bind to H_2O , Cu^{2+} , Zn^{2+} an

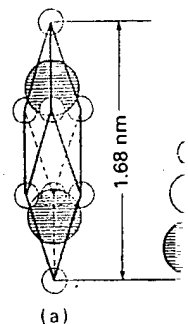


Figure 1 (a) Rhombic pseudorhombohedral

TABLE I Structure

Compound
$\text{MgCl}_2 \cdot 6\text{MeOH}^a$
$\text{MgBr}_2 \cdot 6\text{MeOH}^a$
$\text{NiBr}_2 \cdot 6\text{MeOH}^b$
$\text{CoBr}_2 \cdot 6\text{MeOH}^b$
$\text{CaBr}_2 \cdot 6\text{MeOH}^b$
$\text{CaI}_2 \cdot 6\text{MeOH}^b$
$\text{MgI}_2 \cdot 6\text{MeOH}^b$
$\text{MgCl}_2 \cdot 6\text{EtOH}$
$\text{CaCl}_2 \cdot 4\text{MeOH}^c$
$\text{CaBr}_2 \cdot 4\text{MeOH}^c$
$\text{CoCl}_2 \cdot 4\text{MeOH}^c$
$\text{CaBr}_2 \cdot 4\text{EtOH}$
$\text{NiBr}_2 \cdot 4\text{EtOH}$
$\text{CoBr}_2 \cdot 3\text{EtOH}$
$5\text{MnCl}_2 \cdot 14\text{EtOH}$
$\text{CoCl}_2 \cdot 2.5\text{EtOH}$
$\text{CaCl}_2 \cdot 2\text{MeOH}$
$\text{MgCl}_2 \cdot 2\text{EtOH}$
$\text{HgCl}_2 \cdot 2\text{MeOH}$
$\text{CoCl}_2 \cdot 2\text{MeOH}$
$\text{CoCl}_2 \cdot 2\text{EtOH}$
$\text{NiCl}_2 \cdot 2\text{EtOH}$
$2\text{MnCl}_2 \cdot 3\text{EtOH}$
$\text{NiCl}_2 \cdot \text{MeOH}$

^{a,b,c} Isomorphous

electronegativities (F, O, Cl, N, Br, S, I, C) then it is understood that in various coordinations, chlorine ions may easily displace oxygen atoms. Therefore, other factors, such as steric encumbrance, may play a role in determining the final structural type.

In $\text{MgCl}_2/\text{H}_2\text{O}$ solutions the existence of hexa-aqua complexes of the Mg^{2+} ion has been clearly demonstrated by diffraction measurements [31–37], molecular dynamics [36, 38], neutron inelastic scattering [39], Raman [40] and ^{25}Mg nuclear magnetic resonance (NMR) spectroscopy [41, 42]; the coordination number of eight obtained by Albright [33] is doubtful. Concentrated aqueous solutions of MgCl_2 have a (local) quasi-crystalline structure, a liquid-type quasi-close-packing in the vicinity of the Mg^{2+} cation [34, 37]. In very concentrated solutions of chloride ions in water, hard ions such as Mg^{2+} and Ca^{2+} will bind to H_2O , whereas softer ions such as Ni^{2+} , Cu^{2+} , Zn^{2+} and Cd^{2+} will bind to Cl^- [43]; the

$\text{NiX}_2 \cdot n\text{EtOH}$ systems ($\text{X} = \text{Cl}, \text{Br}$) have been studied to some extent and various compounds were identified ($n = 4, 2$) [44]. In formulating structural hypotheses for the $\text{MgCl}_2 \cdot n\text{EtOH}$ complexes subject of this paper, it is therefore most appropriate to make comparisons to calcium halide solvates, some of which have been the subject of past studies (Table I). In this respect it is noticed that $\text{MgI}_2 \cdot 6\text{MeOH}$ is isomorphous with $\text{CaX}_2 \cdot 6\text{MeOH}$ ($\text{X} = \text{Br}, \text{I}$). Alkyl substituents, such as in alcohols, reduce the hardness in proportion to the size of the alkyl groups. In an ethylalcoholic MgCl_2 solution, chlorine atoms are therefore not expected to belong to the first solvation sphere of magnesium, i.e. $[\text{Mg}(\text{EtOH})_6]^{2+} \cdot 2\text{Cl}^-$. Indeed, in a 1.44 M $\text{MgCl}_2/\text{MeOH}$ solution the cationic solvate shell consists of an octahedral arrangement of OH groups (with the methyl groups having much less tendency to occupy symmetrical positions), [45]. The observed Mg–O distance (0.2068(3) nm) is here somewhat less than usually found in aqueous solution (0.210–0.212 nm) [35], which is physically significant. On the basis of bond valence analysis [46] a value of 0.209 nm is expected. The distance shortening has been ascribed to the different electrostatic interactions between Mg^{2+} and CH_3OH molecules as compared to Mg^{2+} and H_2O due to the difference in the dipole moments in magnitude and in position relative to the lone pair. Raman and ^{25}Mg NMR studies of MgCl_2 in MeOH and MeOH/ H_2O mixtures denote preferential hydration of Mg^{2+} in MeOH– H_2O mixtures; nevertheless, cation–methanol interactions were also detected even with low methanol contents [47, 48]. In the $\text{MgCl}_2\text{--H}_2\text{O--MeOH}$ system, the solid phases existing successively as the methanol content increases are $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$, $\text{MgCl}_2 \cdot 4\text{H}_2\text{O} \cdot 2\text{CH}_3\text{OH}$ and $\text{MgCl}_2 \cdot 6\text{CH}_3\text{OH}$ [49]. In the solid state, the structure

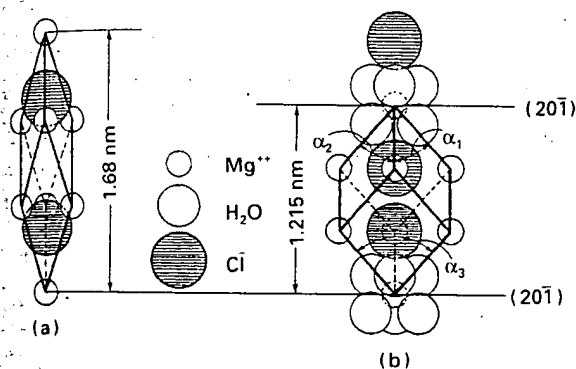


Figure 1 (a) Rhombohedral elementary unit cell of MgCl_2 and (b) pseudorhombic elementary unit cell of $\text{Mg}(\text{H}_2\text{O})_6\text{Cl}_2$ [52].

TABLE I Structural data of $\text{MX}_2 \cdot n\text{ROH}$ complexes

Compound	Structural type	Reference
$\text{MgCl}_2 \cdot 6\text{MeOH}^a$	Discrete octahedral $[\text{M}(\text{MeOH})_6]^{2+}$ groups (ionic)	[53]
$\text{MgBr}_2 \cdot 6\text{MeOH}^a$	Id.	[53]
$\text{NiBr}_2 \cdot 6\text{MeOH}^b$	Id.	[73]
$\text{CoBr}_2 \cdot 6\text{MeOH}^b$	Id.	[53, 73]
$\text{CaBr}_2 \cdot 6\text{MeOH}^b$	Id.	[53, 73]
$\text{CaI}_2 \cdot 6\text{MeOH}^b$	Id.	[53, 74]
$\text{MgI}_2 \cdot 6\text{MeOH}^b$	Id.	[53, 74]
$\text{MgCl}_2 \cdot 6\text{EtOH}$	Discrete octahedral $[\text{M}(\text{EtOH})_6]^{2+}$ groups (ionic)	[75]
$\text{CaCl}_2 \cdot 4\text{MeOH}^c$	Discrete octahedral $\text{M}[\text{transX}_2(\text{ROH})_4]$ groups (molecular)	[62]
$\text{CaBr}_2 \cdot 4\text{MeOH}$	Id.	[55]
$\text{CoCl}_2 \cdot 4\text{MeOH}^c$	Id.	[76]
$\text{CaBr}_2 \cdot 4\text{EtOH}$	Id.	[55]
$\text{NiBr}_2 \cdot 4\text{EtOH}$	Dinuclear $[(\text{EtOH})_4\text{NiBr}_2\text{Ni}(\text{EtOH})_4]^{2+}$ ion	[63]
$\text{CoBr}_2 \cdot 3\text{EtOH}$	$(\text{CoBr}_4)^{2-}$ and $[\text{Co}(\text{EtOH})_6]^{2+}$ ions	[68]
$5\text{MnCl}_2 \cdot 14\text{EtOH}$	Complex polymeric chains MX_2L_4 , MX_4L_2 , MX_3L_3	[77]
$\text{CoCl}_2 \cdot 2.5\text{EtOH}$	$[\text{Co}(\text{EtOH})_6]^{2+}$ and trinuclear $[\text{Cl}_3\text{Co--Cl--Co}(\text{EtOH})_4\text{--Cl--CoCl}_3]^{2-}$ ions	[69]
$\text{CaCl}_2 \cdot 2\text{MeOH}$	$[\text{M}(\text{Cl}_4\text{O}_2)]_n$ chains; doubly chlorine bridged octahedra	[78]
$\text{MgCl}_2 \cdot 2\text{EtOH}$	Id.	[19]
$\text{HgCl}_2 \cdot 2\text{MeOH}$	Id., highly deformed octahedra	[79]
$\text{CoCl}_2 \cdot 2\text{MeOH}$	O_h (IR, VIS, magnetic)	[80]
$\text{CoCl}_2 \cdot 2\text{EtOH}$	C_{3v} (IR, VIS, magnetic)	[80, 81]
$\text{NiCl}_2 \cdot 2\text{EtOH}$	Polynuclear zig-zag chain of $\text{Ni}(\text{Cl}_4\text{O}_2)$ octahedra	[81]
$2\text{MnCl}_2 \cdot 3\text{EtOH}$	Complex polymeric chains MX_4L_2 , MX_3L	[82]
$\text{NiCl}_2 \cdot \text{MeOH}$	Pseudooctahedral	[83]

^{a,b,c} Isomorphous

of $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ can be considered as an arrangement of discrete $\text{Cl}-(\text{H}_2\text{O})_3 \text{Mg}^{2+} (\text{H}_2\text{O})_3\text{Cl}^-$ groups (molecular lattice structure type) [50]; the relation to the MgCl_2 structure is given in Fig. 1. Partin and O'Keeffe [51] have recently discussed the relationship of MgCl_2 to the fluorite structure within the framework of a simple Born-Mayer model. $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ can even more appropriately be compared to a highly deformed fluorite-type lattice with its sizeable $\text{Mg}(\text{H}_2\text{O})_6^{2+}$ pseudocations but with only two instead of four close contacts between each H_2O ligand and Cl^- ions [52]. The isomorphous $\text{MgX}_2 \cdot 6\text{MeOH}$ ($\text{X} = \text{Cl}, \text{Br}$) structures are built up of discrete octahedral $[\text{Mg}(\text{MeOH})_6]^{2+}$ groups; the halogens are again excluded from the coordination sphere of the metal [53]. This is the usual arrangement for hexalcoholates (Table I). If we consider cations replacing the oxygen octahedra around magnesium and anions occupying the tetrahedral interstices, the structure of $\text{MgX}_2 \cdot 6\text{MeOH}$ can be described as a distorted fluorite lattice (though less distorted than in the case of $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$), now with three hydrogen bonds between each halogen and hydroxyl groups [53]. $\text{MgCl}_2 \cdot 6\text{EtOH}$ is not isomorphous with the methanol solvate [54], as indeed would be unexpected (cf. similar absence of isomorphism between $\text{CaBr}_2 \cdot 4\text{MeOH}$ and $\text{CaBr}_2 \cdot 4\text{EtOH}$ [55]). In $\text{MgBr}_2 \cdot 6\text{C}_5\text{H}_5\text{N}$, bromine is included in the magnesium sphere contrary to the case of $\text{MgBr}_2 \cdot 6\text{H}_2\text{O}$ [56]. This is explained by the hard and soft acid concept [29, 30]. Apart from the nature of the metal, factors which determine the structure of an adduct are, therefore, polarity, steric factors and the number and kind (soft/hard) of ligands available.

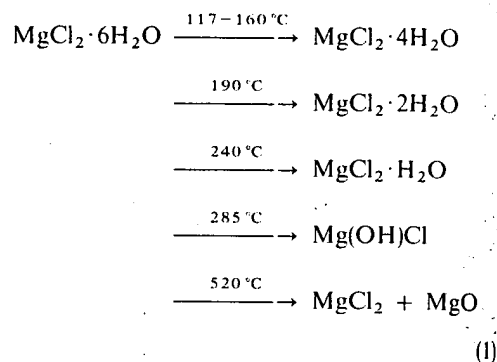
3.2. Dealcoholation

The results of the phase analysis of some 60 preparations of $\text{MgCl}_2 \cdot n\text{EtOH}$ adducts are collected in Table II. X-ray powder spectra could readily be interpreted on the basis of nine characteristic sets of d and I/I_0 values (Table III). It then appears that most of the preparations are admixtures of two adducts with different $\text{EtOH}/\text{MgCl}_2$ ratio. Although this interpretation is fairly straightforward, assignment of the correct stoichiometry to each of the new phases is considerably more difficult. The interpretation given (Table II) is the one which at present best accounts for the observed diffraction data.

The occasional observation of more than two phases in several preparations denotes the absence of equilibrium conditions. As our preparative conditions do not necessarily just lead to stable adducts in the MgCl_2 - EtOH phase diagram, the interpretation of the experimental data might be less reliable in the occurrence of metastable phases. It should also be underscored that the relevance of thermal decomposition data (Fig. 1) to the findings of Table II is limited due to the great differences in preparative conditions. Moreover, our experiments do not necessarily lead to the detection of *all* $\text{MgCl}_2 \cdot n\text{EtOH}$ adducts within the interval ($6 \leq n < 1.4$) studied. Finally, no cases of isomorphism were detected between any of our

$\text{MgCl}_2 \cdot n\text{EtOH}$ adducts with similar solvates which could facilitate structural attributions.

The thermal behaviour of $\text{MgCl}_2 \cdot 6\text{EtOH}$ (at a heating rate of 3°C min^{-1}) shows a steady loss of ethanol with decompositions (peak temperatures) at 92, 120, 163, 168 and 201°C (Fig. 2). At lower heating rates (1 and 2°C min^{-1}) the observed thermal effects indicate an even more complex pattern (Table IV), at variance to the rather smooth degradation suggested by Fig. 1 of [27]. Peak shapes denote the occurrence of multiple processes at various stages (12.4% weight loss corresponds to 1 mol ethanol). The results appear to be more complex than the thermal decomposition of $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$, which proceeds as follows [57]



By general dehydration processes it is possible to remove about four molecules of water without securing undesirable side reactions. However, the evolution of molecular water of crystallization is not the only possible chemical process, and alternative reactions may proceed through intracrystalline interactions of lattice water with other constituents of the salt. This is particularly important upon further dehydration of $\text{MgCl}_2 \cdot 2\text{H}_2\text{O}$ to MgCl_2 , where various undesired side reactions do occur, such as the formation of $\text{Mg}(\text{OH})\text{Cl}$, MgO and HCl [58]. This inconvenience can be avoided using a magnesium chloride entraining solvent (e.g. a higher alcohol) in the extraction-azeotropic zone [2] or by other procedures based on the use of HCl [3], derouting via a magnesium chloride (hex) ammoniate product [2, 6] or those indicated in the experimental section. At reduced pressures ($< 6.666 \text{ kN m}^{-2}$) magnesium basic chloride was not observed by Pribylov and Krasherinnikova [59].

Other stable hydrates, namely $\text{MgCl}_2 \cdot 12\text{H}_2\text{O}$ and $\text{MgCl}_2 \cdot 8\text{H}_2\text{O}$ (in α and β modification) [60] do not find analogues in the alcoholates (Tables I and V) and other adducts. This is not surprising both for steric reasons and in view of the fact that $\text{MgCl}_2 \cdot 6\text{EtOH}$ has a high vapour pressure.

On the basis of the foregoing it should be pointed out that the name "magnesium chloride hydrate" used for $\text{MgCl}_2 \cdot n\text{H}_2\text{O}$ is misleading. These substances are true hydroxides, differing in structure from MgCl_2 (Fig. 1), not species in which water molecules as such have been taken up into the magnesium dichloride lattice, though the chloride has a strong affinity for water and for other polar molecules. The same applies to the "alcohol adducts".

As indicated above, when "hydrates" or "alcoholates" are heated strongly, they lose water or alcohol

TABLE II Phase c

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^a Cf. Section 2.

^b Values given in parentheses.

^c n of $\text{MgCl}_2 \cdot n\text{EtOH}$.

^d After heating of sample.

^e Shows irreversible

^f $[\text{H}_2\text{O}] = 4.47\%$.

TABLE II Phase distribution of $\text{MgCl}_2 \cdot n\text{EtOH}$ preparations

Sample	Experimental method ^a	EtOH/ MgCl_2 ratio ^b	Phase distribution ^c
1	2.5	5.85	6
2	2.5	(6.0)	6
3	2.7	5.25	6 > 4.5
4	2.5	5.25	6
5	2.7	5.15	6 > 4.5
6	2.7	5.15	6 \approx 4.5
7	2.7	(5.0)	6, 4
8	2.7	(5.0)	6, 4
9	2.7	5.0	4.5, some 6
10	2.7	4.87	6, some 3.33
11	2.6	4.80	6, some 4.5
12	2.7	4.75	4.5, some 6
13	2.7	4.56	4.5, some 6
14	2.4	4.0	3.33, some 4.5
15	2.5	3.87	4, some 6, some 3.33
16	2.1	3.60	3.33, some 6, unknown
17	2.1	3.49	3.33
18	2.1	3.27	3.33
19	2.1	3.25	3.33
20	2.1	3.21	3.33
21	2.4	3.20	3.33, 2.5, unknown
22	2.1	3.13	3.33, some 2.5
23	2.1	3.08	3.33, 2.5
24	2.1	3.04	3.33, 2.5
25	2.1	3.00	3.33, some 2.5
26	2.7 + 2.2	(3.00)	3.33, 2.5
27	2.7 + 2.2	(3.00)	3.33, 2.5 (?), some 4
28	2.5	2.90	3.33, 2.5
29	2.1	2.95	3.33, some 2.5
30	2.1	2.86	2.5
31	2.1	2.85	2.5, some 3.33
32	2.1	2.80	2.5
33	2.1	2.68	2.5, some 1.67
34	2.6	2.0 ^f	$\text{MgCl}_2 \cdot 2\text{EtOH} \cdot 0.5\text{H}_2\text{O}$
35	2.1	2.63	2.5, unknown (traces)
36	2.1	2.60	2.5, some 1.67
37	2.1	2.53	2.5, some 1.67
38	2.1	2.51	2.5, some 1.67
39	2.1	2.50	2.5, unknown (traces)
40	2.1	2.49	2.5, some 1.67
41	2.1	2.45	2.5, some 1.67
42	2.1	2.40	2.5, some 1.67, some 1.50
43	2.1	2.36	2.5, some 1.67
44	2.1	2.36	2.5
45	2.1	2.28	2.5, some 1.67
46	2.1	2.27	largely 2.5
47	2.1	2.18	2.5, some 1.5
48 ^d	2.1	~ 2.10	2.5, some unknown
49	2.1	2.09	2.5, some unknown
50 ^e	2.1	2.09	2.5, some 1.67
51	2.1	1.77	1.67
52	2.1	1.72	1.67
53	2.1	1.69	1.67
54	2.1	1.66	1.67
55	2.5 + 2.2	1.60	1.25, amorphous
56	2.1	1.60	1.67
57	2.1	1.58	1.67
58	2.3	1.50	1.5
59	2.1 + 2.2	(3.0 \rightarrow) 1.48	1.25, some unknown
60	2.1 + 2.2	1.39	1.25, amorphous

^a Cf. Section 2.^b Values given in parentheses are nominal compositions.^c n of $\text{MgCl}_2 \cdot n\text{EtOH}$.^d After heating of sample 50 at 85°C for 10 min.^e Shows irreversible exothermal transition at 82°C in nitrogen.^f $[\text{H}_2\text{O}] = 4.47\%$.

TABLE III Interplanar distances of $\text{MgCl}_2 \cdot n\text{EtOH}$ adducts

$\text{MgCl}_2 \cdot 6\text{EtOH}^a$		$\text{MgCl}_2 \cdot 6\text{EtOH}^b$		$\text{MgCl}_2 \cdot 4.5\text{EtOH}^a$		$\text{MgCl}_2 \cdot 4\text{EtOH}^c$		$\text{MgCl}_2 \cdot 3.33\text{EtOH}^a$		$\text{MgCl}_2 \cdot 2.5\text{EtOH}^a$		$\text{MgCl}_2 \cdot 2\text{EtOH} \cdot 0.5\text{H}_2\text{O}^c$		$3\text{MgCl}_2 \cdot 5\text{EtOH}^a$		$\text{MgCl}_2 \cdot 1.5\text{EtOH}^c$		$\text{MgCl}_2 \cdot 1.25\text{EtOH}^c$	
$d(\text{nm})$	(I/I_0)	$d(\text{nm})$	hkl	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0
0.988	> 100	0.995	001	0.918	100	1.016	49	0.966	15	1.403	27	1.119	35	1.449	9	1.046	100	1.163	100
0.674	3	0.688	100	0.776	10	0.876	73	0.937	> 100	0.977	100	1.040	> 100	1.323	100	0.741	48	0.720	25
0.568	5	0.566	101	0.708	24	0.842	100	0.745	15	0.890	70	0.912	26	0.987	50	0.624	46	0.595	8
0.491	39	0.498	002	0.684	32	0.750	58	0.686	27	0.757	11	0.783	100	0.895	33	0.442	12	0.359	8
0.399	26	0.403	102	0.621	20	0.572	42	0.557	11	0.717	6	0.703	11	0.750	8	0.344	17	0.347	16
0.366	22	0.398	110	0.592	6	0.377	21	0.529	15	0.595	6	0.658	12	0.720	33	0.336	14	0.334	8
0.365	30	0.369	111	0.532	19	0.345	79	0.488	19	0.522	14	0.537	21	0.611	13	0.314	38	0.323	7
0.341	5	0.344	200	0.499	13	0.314	17	0.470	18	0.478	10	0.516	24	0.601	4	0.300	68	0.307	16
0.323	4	0.332	003	0.458	11	0.2840	30	0.465	19	0.456	15	0.480	30	0.556	8	0.265	26	0.288	49
0.308	14	0.325	201	0.394	57	0.2522	26	0.456	19	0.432	5	0.442	36	0.528	4	0.255	20	0.278	16
0.2967	53	0.311	112	0.385	33			0.447	19	0.407	8	0.422	36	0.504	12	0.240	19	0.242	28
0.2810	5			0.358	80			0.4031	7	0.392	14	0.394	43	0.456	8	0.223	16	0.232	7
0.2532	4			0.346	19			0.3826	36	0.378	25	0.379	19	0.427	4	0.219	21	0.228	16
0.2475	100			0.343	23			0.3663	100	0.364	11	0.375	12	0.417	6	0.213	18	0.221	9
0.2377	6			0.334	24			0.3520	16	0.353	11	0.363	10	0.399	4	0.206	16	0.190	11
0.2334	20			0.324	27			0.3401	44	0.334	26	0.360	19	0.394	8	0.181	12		
0.2330	25			0.312	20			0.3317	16	0.319	18	0.351	30	0.380	8				
0.2300	8			0.309	13			0.3226	39	0.313	8	0.345	24	0.355	6				
0.2100	1			0.300	19			0.3081	27	0.2998	14	0.336	16	0.345	6				
0.2038	10			0.296	11			0.2912	24	0.2950	5	0.323	31	0.340	12				
0.2010	6			0.2865	27			0.2879	20	0.2891	11	0.314	31	0.335	10				
0.1985	7			0.2650	8			0.2788	51	0.2770	10	0.308	33	0.322	20				
0.1907	5			0.2539	15			0.2714	26	0.2736	22	0.303	26	0.311	22				
0.1792	9			0.2383	9			0.2679	46	0.2684	20	0.294	8	0.300	3				
0.1774	7			0.2356	10			0.2484	34	0.2632	5	0.276	23	0.2940	17				
0.1720	3			0.2278	10			0.2471	8	0.2569	16	0.267	43	0.2907	25				
0.1606	5			0.2145	9			0.2442	5	0.2471	8	0.260	15	0.2872	26				
0.1578	6			0.2115	6			0.2404	8	0.2442	5	0.253	15	0.2814	30				
				0.2015	7			0.2383	11	0.2404	8	0.247	20	0.2790	12				
				0.1984	6					0.2383	11	0.245	39	0.2741	9				
												0.219	25	0.2692	10				
												0.197	6	0.2550	6				
												0.190	11	0.2456	18				
														0.2420	5				
														0.2361	10				
														0.2328	7				
														0.2222	15				
														0.2142	3				
														0.2101	4				
														0.2083	5				

TABLE III (Continued)

$\text{MgCl}_2 \cdot 6\text{EtOH}^a$		$\text{MgCl}_2 \cdot 6\text{EtOH}^b$		$\text{MgCl}_2 \cdot 4.5\text{EtOH}^a$		$\text{MgCl}_2 \cdot 4\text{EtOH}^c$		$\text{MgCl}_2 \cdot 3.33\text{EtOH}^a$		$\text{MgCl}_2 \cdot 2.5\text{EtOH}^a$		$\text{MgCl}_2 \cdot 2\text{EtOH} \cdot 0.5\text{H}_2\text{O}^c$		$3\text{MgCl}_2 \cdot 5\text{EtOH}^a$		$\text{MgCl}_2 \cdot 1.5\text{EtOH}^c$		$\text{MgCl}_2 \cdot 1.25\text{EtOH}^c$	
$d(\text{nm})$	(I/I_0)	$d(\text{nm})$	hkl	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0	$d(\text{nm})$	I/I_0

0.2328 7
0.2222 15
0.2142 3
0.2101 4
0.2083 5

TABLE III (Continued)

MgCl ₂ ·6EtOH ^a		MgCl ₂ ·4.5EtOH ^a		MgCl ₂ ·4EtOH ^a		MgCl ₂ ·3.33EtOH ^a		MgCl ₂ ·2.5EtOH ^a		MgCl ₂ ·2EtOH·0.5H ₂ O ^a		MgCl ₂ ·1.5EtOH ^a		MgCl ₂ ·1.25EtOH ^a	
d(nm)	(I/I ₀)	d(nm)	I/I ₀	d(nm)	I/I ₀	d(nm)	I/I ₀	d(nm)	I/I ₀	d(nm)	I/I ₀	d(nm)	I/I ₀	d(nm)	I/I ₀
0.2038	5	0.2038	5	0.2038	5	0.2038	5	0.2038	5	0.2038	5	0.2038	5	0.2038	5
0.1979	5	0.1979	5	0.1979	5	0.1979	5	0.1979	5	0.1979	5	0.1979	5	0.1979	5
0.1903	8	0.1903	8	0.1903	8	0.1903	8	0.1903	8	0.1903	8	0.1903	8	0.1903	8
0.1831	4	0.1831	4	0.1831	4	0.1831	4	0.1831	4	0.1831	4	0.1831	4	0.1831	4
0.1807	9	0.1807	9	0.1807	9	0.1807	9	0.1807	9	0.1807	9	0.1807	9	0.1807	9
0.1794	8	0.1794	8	0.1794	8	0.1794	8	0.1794	8	0.1794	8	0.1794	8	0.1794	8
0.1705	5	0.1705	5	0.1705	5	0.1705	5	0.1705	5	0.1705	5	0.1705	5	0.1705	5
0.1687	8	0.1687	8	0.1687	8	0.1687	8	0.1687	8	0.1687	8	0.1687	8	0.1687	8

^a Recorded at 1/4° 20 min⁻¹ chart speed.

^b Calculated on the basis of trigonal space group P3, a = 0.795 nm, c = 0.995 nm, atomic parameters from [75].

^c Recorded at 1° 20 min⁻¹ chart speed.

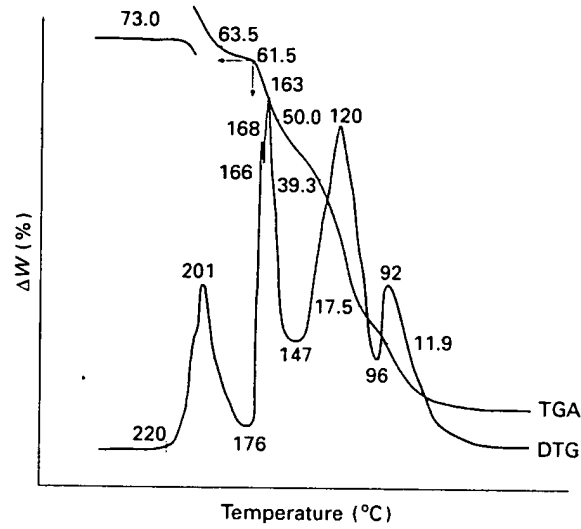


Figure 2 Thermal decomposition of MgCl₂·6EtOH in a nitrogen atmosphere at a heating rate of 3°C min⁻¹.

and are converted ultimately into the chloride, MgCl₂, by way of a series of intermediate phases recognizable as structurally distinct by means of their X-ray powder diffraction patterns. The intermediate phases encountered in a particular (de)solvation experiment depend critically on the starting material and experimental conditions used [11, 22]. In addition to our experimental techniques, adducts have also been prepared by grinding a mixture of magnesium dihalide with the desired amount of organic electron-donor compound or by spray-drying [11]. Also treatment of MgCl₂-alcohol complexes with SiCl₄ or organoaluminium compounds results in partial removal of alcohol, causing further collapse of the MgCl₂ crystal aggregates [11, 12]. Slow elimination of ethanol from MgCl₂·6EtOH at 50°C and 0.1 atm (i.e. in conditions close to our experimental procedure in Section 2.4) has recently been reported over the full composition range [20]. For 2 < n ≤ 6 ionic Mg-Cl bonds prevail in the structure of the adducts. For n ≤ 2, covalent bonds are observed among chlorine and magnesium atoms, which give rise to chain compounds through chlorine bridges among the magnesium atoms. This conforms schematically to a process MgCl₂·nB → MgCl₂·2B → activated MgCl₂, as observed for many electron donors.

In our experimental conditions the most frequently encountered phases of dealcoholation of MgCl₂·6EtOH are MgCl₂·3.33EtOH, MgCl₂·2.5EtOH and MgCl₂·1.67EtOH. It is interesting to notice the absence of MgCl₂·3EtOH, a jet-crystallized product mentioned by Neste Oy (61) and of MgCl₂·2EtOH, described by Di Noto *et al.* [19].

(Sub)solvated dimetal halogenides exhibit a great variety of structural types with variation in stereochemical metal coordination and the formation of either discrete structure units or condensed species (from short to infinite polymeric chains). Essentially three structural types may be distinguished:

(a) structures composed of only one discrete mononuclear species, either ionic (when charged) as in

TABLE IV TGA and DTA of $\text{MgCl}_2 \cdot 6\text{EtOH}$ in a nitrogen atmosphere

Heating rate 1°C min^{-1}		2°C min^{-1}		3°C min^{-1}	
T interval ($^\circ\text{C}$)	$-\Delta W$ (%)	T interval ($^\circ\text{C}$)	$-\Delta W$ (%)	T interval ($^\circ\text{C}$)	$-\Delta W$ (%)
25–35	2.96 ^a	25–35	1.75	25–90	11.9
35–82	13.4	35–105	20.55	90–107	5.6
82–102	20.4	105–133	14.34	107–131	21.8
102–105	2.08	133–158	12.4	131–158	10.7
105–123	3.51	158–163	12.3	158–169	11.5
123–128	3.4	163–195	3.58	169–192	2.0
128–131	0.99	195–220	7.5	192–210	9.5
131–134	1.42		$\Delta W_{\text{tot}} = -72.4\%$		$\Delta W_{\text{tot}} = -73.0\%$
134–140	4.06				
140–150	4.5				
150–156	4.5				
156–193	8.3				
193–210	2.19				
	$\Delta W_{\text{tot}} = -71.7\%$				

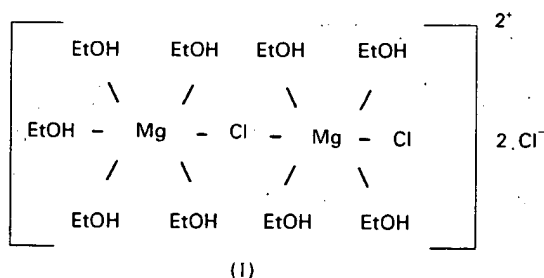
^a Weight losses are referred to inflection points of TG curve.

$\text{CoBr}_2 \cdot 6\text{MeOH}$ and $\text{MgCl}_2 \cdot 6\text{EtOH}$, or molecular (uncharged) as in $\text{CoCl}_2 \cdot 4\text{MeOH}$ and $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$;

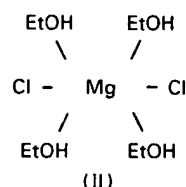
(b) polymeric polynuclear structures, such as $\text{NiBr}_2 \cdot 4\text{EtOH}$, $\text{CaCl}_2 \cdot 2\text{MeOH}$, $\text{HgCl}_2 \cdot 2\text{MeOH}$ and $2\text{MnCl}_2 \cdot 3\text{EtOH}$, resulting from one- to three-dimensional association of octahedra;

(c) intermediate structures with complex mononuclear and/or polynuclear species (e.g. $\text{CoCl}_2 \cdot 2.5\text{EtOH}$ and $\text{CoBr}_2 \cdot 3\text{EtOH}$).

As may be seen from Table II, two solvates are observed in the intermediate range of $\text{MgCl}_2 \cdot 6\text{EtOH}$ and $\text{MgCl}_2 \cdot 3.33\text{EtOH}$. One of these is present in the range $5.25 < n < 4.0$ and is provisionally assigned as $\text{MgCl}_2 \cdot 4.75\text{EtOH}$ or $\text{MgCl}_2 \cdot 4.5\text{EtOH}$. The latter compound could display the following structure (I)

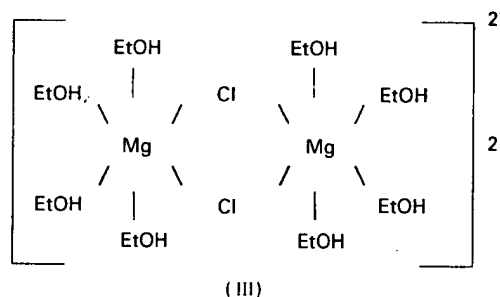


The second species was occasionally present in the $5.0 < n < 3.0$ range and, like $\text{MgCl}_2 \cdot 4.5\text{EtOH}$, was never found as a single-phase product but always in admixture. The best possible hypothesis is $\text{MgCl}_2 \cdot 4\text{EtOH}$, which is most simply represented as (II)



with *trans* arrangement of the halides in a deformed octahedral structure, in accordance with other metal halides solvated with four molecules of a monodentate

solvent (Table I); usually, the O_h symmetry around the metal is slightly deformed [62]. Alternatively, the existence of a binuclear ion (III)



could also be admitted, as found in $\text{NiBr}_2 \cdot 4\text{EtOH}$ [63] or in $\alpha\text{-CaCl}_2 \cdot 4\text{H}_2\text{O}$ [64] though with a different structural unit. Infinite chains of doubly halogen bridged $\text{Cd}(4\text{X}, 2\text{H}_2\text{O})$ octahedra are reported in $\text{CdX}_2 \cdot 4\text{H}_2\text{O}$ ($\text{X} = \text{Cl}, \text{Br}$), which are actually both better described as $[\text{CdX}_2 \cdot 2\text{H}_2\text{O}] \cdot 2\text{H}_2\text{O}$, i.e. hydrated $\text{CdX}_2 \cdot 2\text{H}_2\text{O}$ with two water molecules by insertion [65, 66]. The latter examples illustrate that EtOH in $\text{MgCl}_2 \cdot n\text{EtOH}$ may, in principle, be either chemically bound as a ligand or physically built into the crystal lattice as a true solvate. $\text{MgBr}_2 \cdot 6\text{C}_5\text{H}_5\text{N}$ is another example in case [56].

$\text{MgCl}_2 \cdot 3.33\text{EtOH}$ or $3\text{MgCl}_2 \cdot 10\text{EtOH}$ is observed for $4.0 > \text{EtOH}/\text{MgCl}_2 > 2.9$ and appears as the only phase in samples with $n = 3.21, 3.25, 3.27$ and 3.49 (Table II). Reasonable structural hypotheses for this phase are ionic or molecular trinuclear structures such as IV or V

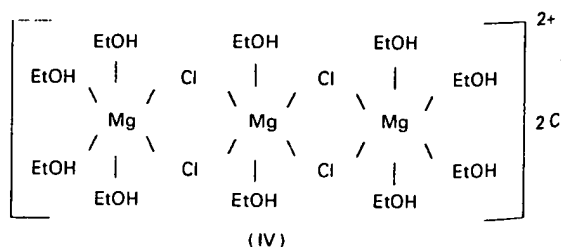
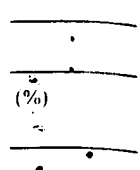


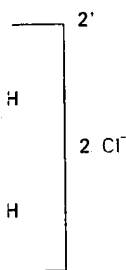
TABLE V Structural data of addition products of magnesium halides

Adduct ^a	Ligand set	Coordination	Bond distances (nm)	Reference
$\text{MgCl}_2 \cdot n\text{H}_2\text{O}$	$\text{O}(\text{H}_2\text{O})$	Octahedral	$\text{Mg}-\text{O}$	[62]
$\text{MgCl}_2 \cdot n\text{EtOH}$	$\text{O}(\text{H}_2\text{O})$	Octahedral	$\text{Mg}-\text{O}$	[62]



73.0%

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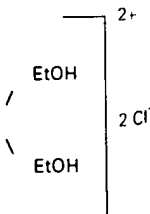


TABLE V Structural data of addition products of magnesium halides

Adduct ^a	Ligand set	Coordination	Bond distances (nm)		Reference
			Mg-X	Mg-O	
$\text{MgCl}_2 \cdot 12\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	[0.2059, 0.2061, 0.2065 (7)] 2x	[84]
$\text{MgCl}_2 \cdot 2\text{CdCl}_2 \cdot 12\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	[0.2052 (5)] 4x, [0.2081 (4)] 2x	[85]
$2\text{MgCl}_2 \cdot \text{CdCl}_2 \cdot 12\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	[0.2041, 0.2096 (4)] 3x; [0.2071 (4)] 6x	[86]
$\text{MgCl}_2 \cdot 10\text{H}_2\text{O} \cdot 2(\text{CH}_3)_6\text{N}_4$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.210 (-) 6x	[87]
$\text{MgCl}_2 \cdot 6\text{H}_2\text{O} \cdot \text{dx}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	[0.2057, 0.2060, 0.2073 (4)] 2x	[88]
$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}^{\text{b,c}}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	[0.2057, 0.2062 (1)] 3x	[89]
$\text{MgClBr} \cdot 6\text{H}_2\text{O}^{\text{c}}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	n.d.	[50]
$\text{MgBr}_2 \cdot 6\text{H}_2\text{O}^{\text{c}}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	n.d.	[89]
$\text{MgCl}_2 \cdot \text{KCl} \cdot 6\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.2028, 0.2042, 0.2042, 0.2045, 0.2053, 0.2053 (2); 0.2040, 0.2041, 0.2042, 0.2049, 0.2050, 0.2056 (2)	[52]
$\text{MgCl}_2 \cdot \text{RbCl} \cdot 6\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.1991, 0.2020, 0.2062, 0.2071, 0.2078, 0.2115 (14); 0.1982, 0.2007, 0.2016, 0.2027, 0.2088, 0.2097 (14)	[90]
$\text{MgCl}_2 \cdot \text{CsCl} \cdot 6\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.2035, 0.2041, 0.2062, 0.2062, 0.2076, 0.2080 (9)	[91]
$\text{MgCl}_2 \cdot \text{NH}_4\text{Cl} \cdot 6\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.198, 0.200, 0.202, 0.210, 0.210, 0.213 (1); 0.196, 0.199, 0.204, 0.209, 0.212, 0.214 (1)	[92]
$\text{MgX}_2 \cdot 2\text{HgX}_2 \cdot 6\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.2019, 0.2038 (2x), 0.2087, 0.2139 (26) (2x)	[93]
$\text{MgCl}_2 \cdot \text{toc} \cdot 6\text{H}_2\text{O}$	$\text{O}_6(\text{H}_2\text{O})$	Octahedral	-	0.2054, 0.2057, 0.2059 (2x), 0.2082 (2) (2x)	[94]
$\text{MgCl}_2 \cdot 6\text{C}_2\text{H}_5\text{OH}$	O_6	Octahedral	-	0.2069 (3) 6x	[95]
$\text{MgX}_2 \cdot 6\text{CH}_3\text{OH}$ (X = Cl, Br) ^e	O_6	Octahedral	-	(Br) 0.207 (1) 6x	[75]
$[\text{Mg}(\text{CH}_3\text{COOC}_2\text{H}_5)_6][\text{AlCl}_4]_2$	O_6	Octahedral	-	0.2053 (4) 6x	[53, 74]
$\text{MgBr}_2 \cdot 4\text{THF} \cdot 2\text{H}_2\text{O}$	O_6	Octahedral	-	[0.2042, 0.2117, 0.2164 (7)] (2x)	[96]
$\text{MgCl}_2 \cdot \text{C}_6\text{H}_{11}\text{O}_6 \cdot 4\text{H}_2\text{O}$	O_6	Octahedral	-	0.2032, 0.2038, 0.2051, 0.2079, 0.2094, 0.2143 (4)	[97]
$[\text{MgBr}_2 \cdot 4\text{C}_5\text{H}_5\text{N}] \cdot 2\text{C}_5\text{H}_5\text{N}$	$\text{O}_6(4\text{H}_2\text{O} + 2\text{OH})$	Octahedral	-	-	[98]
$\text{MgX}_2 \cdot 4\text{THF}$ (X = Cl, Br)	<i>trans</i> Br_2N_4 <i>trans</i> X_2O_4	Octahedral	[0.2767 (8)] 2x	-	[56]
$\text{MgBr}_2 \cdot 3\text{THF}^{\text{f}}$	Br_2O_3	Octahedral	Br: [0.2625 (25)] 2x	0.2165 (15) 4x	[14, 97, 99]
$\text{MgX}_2 \cdot 2\text{THF}$ (X = Cl, Br)	$\text{X}_2\text{O}_2^{\text{g}}$	Trigonal bipyramid	-	-	[97]
$\text{MgBr}_2 \cdot 2(\text{CH}_3\text{CH}_2)_2\text{O}$	Br_2O_2	bipyramid	Br: [0.2633, 0.2799 (14)] 2x	0.2126 (13) 2x	[14, 97]
$\text{MgCl}_2 \cdot 2\text{HCOOC}_2\text{H}_5$	Cl_4O_2	Octahedral	0.318, 0.332 (-) 2x	0.213, 0.216 (-) 2x	[100]
$\text{MgCl}_2[\text{CH}_3\text{COOC}_2\text{H}_5]_2 \cdot 1/2\text{CH}_3\text{COOC}_2\text{H}_5$	Cl_4O_2	Tetrahedral	0.274, 0.302 (-) 2x	0.209, 0.214 (-) 2x	[72]
	Cl_4O_2	Octahedral	[0.2494 (10)] 4x	[0.2077 (20)] 2x	[71]
		Octahedral	0.2477, 0.2496, 0.2515, 0.2516 (2)	0.2101, 0.2107 (3)	

^a dx, dioxan; THF, tetrahydrofuran; toc, 1,4,7,10-tetraoxacyclodecane.

^b Bischofite, ND.

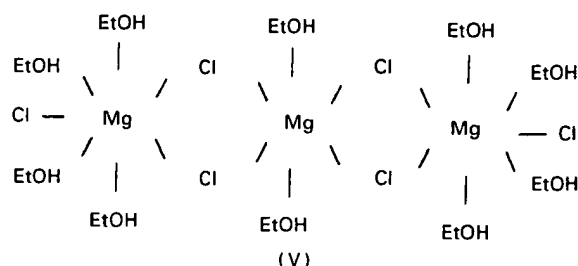
^c Isomorphous.

^d Random occupancy of anion sites by chlorine and bromine atoms.

^e $2\text{H}_2\text{O}$ (*trans*) + 4THF.

^f Only powder data.

^g Polymeric structures with two *trans* THF groups, two interformula unit *cis* Mg-X bonds and two intraformula unit *cis* Mg-X bonds.



The second structural type with triangular prisms linked by two Cl bridges is found (in dimeric form) in α - $\text{CaCl}_2 \cdot 4\text{H}_2\text{O}$ [64]. A seven coordination around magnesium is less likely, the reason why the first option is preferred. We did not find evidence for the existence of $\text{MgCl}_2 \cdot 3\text{EtOH}$. Whereas the corresponding CaBr_2 adduct has been reported, though without much detail [67], it appears that only the crystal structure of one triethanolate ($\text{CoBr}_2 \cdot 3\text{EtOH}$) has been determined [68].

In the range of low ethanol contents, $\text{MgCl}_2 \cdot 2.5\text{EtOH}$ is found extensively for $3.20 < n < 2.09$ with excess of $\text{MgCl}_2 \cdot 3.33\text{EtOH}$ at high EtOH/ MgCl_2 ratios and of $\text{MgCl}_2 \cdot 1.67\text{EtOH}$ at low ratios. The X-ray powder spectrum (Table III) is recognizable in the bar diagram of an unidentified $\text{MgCl}_2 \cdot n\text{EtOH}$ adduct in the $3.08 \leq n \leq 2.24$ range (Fig. 4 of [20]), prepared according to a method closely similar to our experimental procedure in Section 2.4. On the basis of the only known compounds with a similar monodentate solvate/ MX_2 ratio, $\text{CoCl}_2 \cdot 2.5\text{EtOH}$ and $\text{CdCl}_2 \cdot 2.5\text{H}_2\text{O}$, it is difficult to make any structural assignment. $\text{MgCl}_2 \cdot 2.5\text{EtOH}$ does not seem to be isomorphous with $\text{CoCl}_2 \cdot 2.5\text{EtOH}$, which is composed of trinuclear $[\text{Cl}_3\text{Co}-\text{Cl}-\text{Co}(\text{EtOH})_4-\text{Cl}-\text{CoCl}_3]^{2-}$ and $[\text{Co}(\text{EtOH})_6]^{2+}$ ions [69]. In $\text{CdCl}_2 \cdot 2.5\text{H}_2\text{O}$, one H_2O molecule is physically bound as crystal water; the structure exhibits a complicated framework structure [70].

$\text{MX}_2 \cdot 2\text{ROH}$ adducts have been found in various $\text{MX}_2 \cdot \text{ROH}$ systems (Tables I and VI), but we failed to observe this stoichiometry in our experiments. These

compounds are usually built up of infinite planar chains of doubly halogen bridged octahedra with a $4X, 2O$ coordination, i.e. $[\text{MgCl}_2(\text{EtOH})_x]_n$, as also suggested for $\text{MgCl}_2 \cdot 2\text{EtOH}$ [19]. $\text{MgCl}_2 \cdot 2$ ethylacetate [71] and $\text{MgCl}_2 \cdot 2\text{EB}$ [18] exhibit a similar structure at variance to $\text{MgCl}_2 \cdot 2$ ethylformate [72].

It is of interest to notice that Di Noto *et al.* [19] report that $\text{MgCl}_2 \cdot 2\text{B}$ adducts may be obtained by direct synthesis, except for $\text{MgCl}_2 \cdot 2\text{EtOH}$ for which slow elimination of alcohol from $\text{MgCl}_2 \cdot 6\text{EtOH}$ is necessary (reflux time of 20 min at 150°C). Fig. 4 of [20] and Table II of the present study suggest that the experimental methods play a crucial role with regard to the nature of the reaction product(s) obtained.

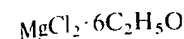
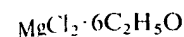
Further dealcoholation in the approximate EtOH/ MgCl_2 range of $2.0 < n < 1.5$ leads to a compound which is best described as $\text{MgCl}_2 \cdot 1.67\text{EtOH}$ or $3\text{MgCl}_2 \cdot 5\text{EtOH}$. At even lower analytical EtOH/ MgCl_2 ratios (approximately 1.5) our data are scarce but X-ray diffraction spectra indicate that at least two more compounds are formed. One of the spectra appears to correspond rather closely to the poorly resolved X-ray diagrams of $\text{MgCl}_2 \cdot n\text{EtOH}$ ($2.10 \leq n \leq 1.06$) of [20]. The stoichiometry $\text{MgCl}_2 \cdot 1.25\text{EtOH}$ is provisionally assigned as being best in accordance with the now available evidence.

As opposed to work by Di Noto *et al.* [20] dealcoholation below EtOH/ $\text{MgCl}_2 \approx 1.5$ was not investigated systematically in this study but was observed to lead to δ - MgCl_2 or more complicated structures, in agreement with this reference.

Summarizing the results, we obtain the following (sequence of) phases: $\text{MgCl}_2 \cdot 6\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 4.5\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 4\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 3.33\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 2.5\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 1.67\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 1.5\text{EtOH} \rightarrow \text{MgCl}_2 \cdot 1.25\text{EtOH} \rightarrow \delta\text{MgCl}_2$, whereas $\text{MgCl}_2 \cdot 2\text{EtOH}$ was not observed in this study; the existence of $\text{MgCl}_2 \cdot 3\text{EtOH}$, as a well-defined compound, has not been proven.

Using quite a different desolvation route (heating in argon in a glass tube at 250 – 400°C for 2.5–6 h)

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4. Conclusio

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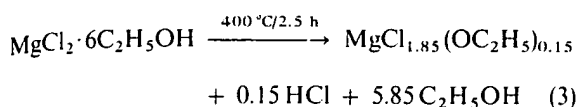
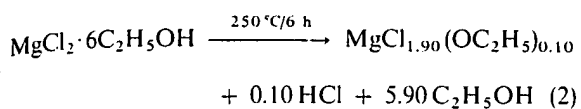
TABLE VI Structural data of some sub-solvate MX_2 adducts.

Adduct	Coordination	Structural type	Reference
$\alpha\text{CaCl}_2 \cdot 4\text{H}_2\text{O}$	Triangular prismatic	Dimeric $\text{Ca}(\text{Cl}_3\text{O}_4)_2$; double chlorine bridge	[64]
$\text{MgCl}_2 \cdot 4\text{H}_2\text{O}$	D_{4h}	$\text{Mg}(\text{Cl}_2\text{O}_4)$	[101]
$\text{MnCl}_2 \cdot 4\text{H}_2\text{O}^a$	D_{4h}	$\text{Mn}(\text{Cl}_2\text{O}_4)$	[102]
$\text{MnBr}_2 \cdot 4\text{H}_2\text{O}^a$	D_{4h}	$\text{Mn}(\text{Br}_2\text{O}_4)$	[103, 104]
$[\text{CdCl}_2 \cdot 2\text{H}_2\text{O}] \cdot 2\text{H}_2\text{O}^b$	D_{2h}	$\text{M}(\text{X}_4\text{O}_2)$ infinite chains; double halogen bridge	[66]
$[\text{CdBr}_2 \cdot 2\text{H}_2\text{O}] \cdot 2\text{H}_2\text{O}^b$	D_{2h}	<i>id.</i>	[65]
$[\text{CdCl}_2 \cdot 1.5\text{H}_2\text{O}] \cdot \text{H}_2\text{O}$	MX_4L_2 , MX_5L	$\text{Cd}_4(\text{Cl}_5\text{O})_2(\text{Cl}_4\text{O}_2)_2$ framework	[70]
$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$	D_{2h}	$\text{M}(\text{X}_4\text{O}_2)$ infinite chains; double halogen bridge	[105]
$\text{CdCl}_2 \cdot 2\text{HCONH}_2$	D_{2h}	<i>id.</i>	[106]
$\text{MgCl}_2 \cdot 2\text{H}_2\text{O}$	D_{2h}	<i>id.</i>	[101]
$\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$	Coplanar	$\text{Cu}(\text{Cl}_2\text{O}_2)$	[107, 108]
$\text{MnCl}_2 \cdot 2\text{H}_2\text{O}$, $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}^a$	D_{2h}	$\text{M}(\text{Cl}_4\text{O}_2)$	[109–111]
$\text{CdCl}_2 \cdot \text{H}_2\text{O}$	Octahedral	Double $\text{Cd}(\text{Cl}_5\text{O})$ chains	[112]

^a Isostructural

^b Not isostructural

Taveira Magalhães *et al.* [24] have given evidence for the following processes



Comparison of Fig. 2 of [24] and Table II shows that no similar alkoxymagnesium chlorides were formed in our case. In analogy to [24], Turova *et al.* [27] established that heating of $\text{MgCl}_2 \cdot 6\text{ROH}$ in the temperature range 20–350°C yields products of the composition $\text{MgCl}_{1.75}(\text{OMe})_{0.25}$ and $\text{MgCl}_{1.80}(\text{OEt})_{0.20}$. Experimental reaction conditions play a major role in the nature of the decomposition process. These results are complementary to the experiments described in this paper.

4. Conclusion

The carefully controlled decomposition of $\text{MgCl}_2 \cdot n\text{EtOH}$ adducts has allowed us to establish the influence of various CTEM process parameters (such as time, pressure, temperatures, concentration, dimensions of the ejection tube, $\text{EtOH}/\text{MgCl}_2$ ratio, presence of ethylbenzoate, nature of the quenching medium, etc.) on the morphology and consistency of the microsphere particles, suitable as catalyst supports in gas-phase Ziegler–Natta α -olefin polymerization.

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Influence and paper a coated

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Papers coated with a mechanical layer performed to paper on the from measurement height of tar separate layer influence of the interaction the experimental calculation explained as interaction of coated paper X-ray analysis

1 Introduction

A coated paper is a paper sheet with a polymer coating. A smooth surface and surface gloss are proposed to differ in printing and printing. They investigate their dynamic mechanical analysis study of such modes of deformation, have different the specimen properties of it. Studies of fibre an example of dynamic mechanical. Because dynamic is sensitive to small differences in structure with secondary temperatures, this differences in properties. In the coating developed between